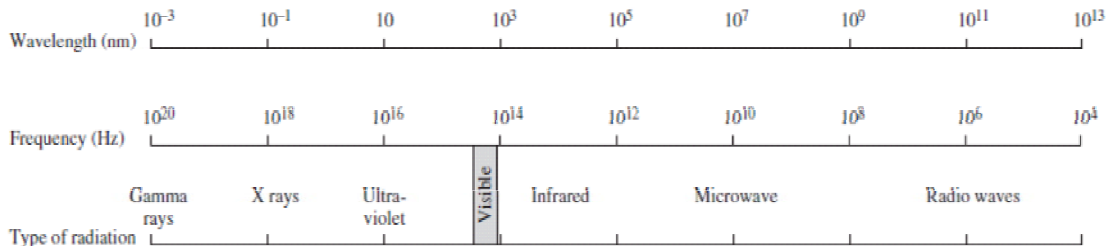


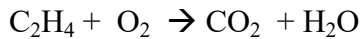




\_\_\_\_ 4. Using the figure below, which radiation has the lowest frequency?

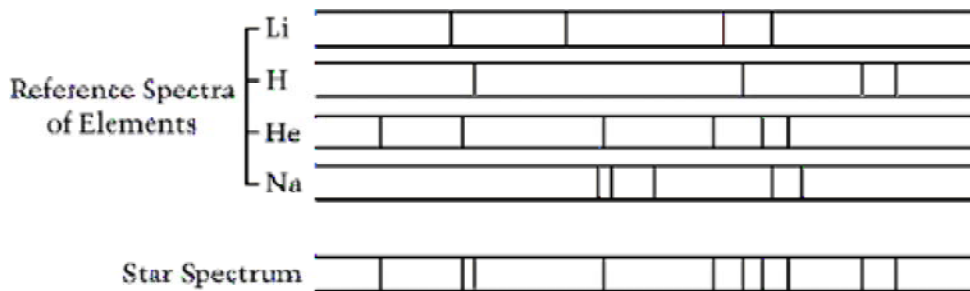


- a. X rays  
 b. Ultraviolet  
 c. Microwave  
 d. Gamma rays
- \_\_\_\_ 5. The correct coefficient for oxygen when the equation is balanced is:



- a. 1  
 b. 2  
 c. 3  
 d. 6  
 e. 14
- \_\_\_\_ 6. Choose the correct valence electrons for lead.
- a.  $6s^2 4f^{14} 5d^{10} 6p^2$   
 b.  $7s^2 7p^2$   
 c.  $6s^2 6p^2$   
 d.  $5s^2 4d^{10} 5p^2$
- \_\_\_\_ 7. When a metallic oxide, like KOH, reacts with water the product is a(an):
- a. acid  
 b. base
- \_\_\_\_ 8. Which element(s) makeup the star spectrum below.

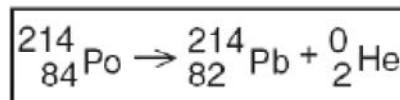
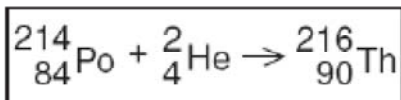
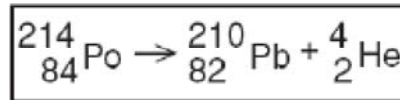
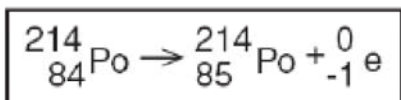
#### EMISSION SPECTRA



- a. He and Na  
 b. He and Li  
 c. He and H  
 d. Li and H
- \_\_\_\_ 9. Choose the correct formula for Ammonium oxalate.
- a.  $(\text{NH}_4)_2\text{C}_2\text{H}_3\text{O}_2$   
 b.  $\text{NH}_4\text{C}_2\text{O}_4$   
 c.  $(\text{NH}_4)_2\text{C}_2\text{O}_4$   
 d.  $\text{C}_2\text{O}_4(\text{NH}_4)_2$
- \_\_\_\_ 10. Which of the following is **NOT** a physical change?
- a. Rusting  
 b. Slicing  
 c. Breaking  
 d. Melting



\_\_\_ 18. Which equation correctly represents the alpha decay of Polonium-214



\_\_\_ 19. What is the polarity of SiO<sub>2</sub>?

a. ionic

c. polar

b. nonpolar

\_\_\_ 20.  $\text{LiOH} + \text{H}_3\text{PO}_4 \rightarrow \text{Li}_3\text{PO}_4 + \text{H}_2\text{O}$

Classify the type of reaction.

a. combustion

c. single replacement

b. neutralization

d. double replacement

\_\_\_ 21. Choose the correct name for SnS<sub>2</sub>

a. Tin sulfide

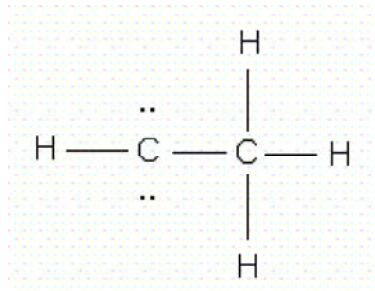
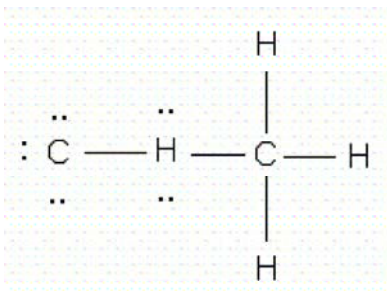
c. Tin (IV) sulfide

b. Tin disulfide

d. Tin (II) sulfide

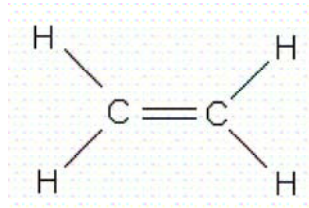
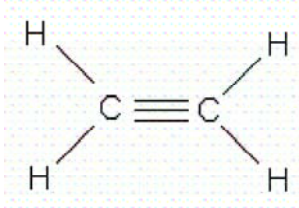
\_\_\_ 22.

Which of the following is the correct Lewis structure for ethylene, C<sub>2</sub>H<sub>4</sub>?



a.

c.



b.

d.

\_\_\_ 23. This piece of equipment is used to heat small amount of substances at high temperatures.

a. hot plate

c. crucible and cover

b. beaker

d. graduated cylinder

\_\_\_ 24. What is the correct name for the following N<sub>2</sub>O<sub>4</sub>?

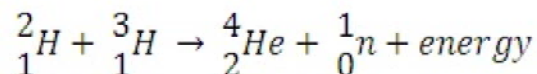
a. nitrogen tetraoxide

c. dinitrogen oxide

b. nitrogen (IV) oxide

d. dinitrogen tetroxide

\_\_\_\_\_ 25. The following equation is an example of:



- \_\_\_\_\_ 26. Which of the following elements can exist as a diatomic molecule?
- |       |       |
|-------|-------|
| a. Br | c. P  |
| b. Se | d. Se |

### Multiple Response

Identify one or more choices that best complete the statement or answer the question.

- \_\_\_\_\_ 27. Which of the following are mixtures?
- |              |         |
|--------------|---------|
| a. Gatorade  | c. Salt |
| b. Diet Coke | d. Iron |
- \_\_\_\_\_ 28. The following compound, NaOH, can be classified as:
- |                                 |                   |
|---------------------------------|-------------------|
| a. does not conduct electricity | d. ionic compound |
| b. molecule                     | e. formula unit   |
| c. does conduct electricity     |                   |
- \_\_\_\_\_ 29. The following compound, C<sub>2</sub>H<sub>2</sub>, can be classified as:
- |                             |                                 |
|-----------------------------|---------------------------------|
| a. does conduct electricity | c. ionic compound               |
| b. molecule                 | d. does not conduct electricity |
- \_\_\_\_\_ 30. What intermolecular forces are present in CH<sub>3</sub>OH?
- |                  |                     |
|------------------|---------------------|
| a. Dipole-Dipole | c. Hydrogen Bonding |
| b. Dispersion    | d. Ionic Bonding    |
- \_\_\_\_\_ 31. Which groups below can be classified as representative elements?
- |                |                         |
|----------------|-------------------------|
| a. halogens    | c. all group B elements |
| b. noble gases | d. alkali earth metals  |
- \_\_\_\_\_ 32. Covalent compounds are:
- |   |  |
|---|--|
| a. Electrons are shared.                                | d. Held together by intermolecular forces. |
| b. Held together by electrostatic forces.               | e. Electrons are transferred.              |
| c. At room temperature they are a solid, liquid or gas. |  |
- \_\_\_\_\_ 33. Ionic compounds are:
- |   |  |
|---|--|
| a. Electrons are shared.                                | d. Held together by intermolecular forces. |
| b. Held together by electrostatic forces.               | e. Electrons are transferred.              |
| c. At room temperature they are a solid, liquid or gas. | f. Formula units                           |
- \_\_\_\_\_ 34. Which of the following compounds are polar?
- |                                  |                     |
|----------------------------------|---------------------|
| a. CH <sub>3</sub> OH            | c. HCN              |
| b. C <sub>2</sub> H <sub>2</sub> | d. SiO <sub>2</sub> |

Name: \_\_\_\_\_

ID: G

### Matching

- a. chemical burn
- b. irritant
- c. thermal burn

- d. forceps
- e. beaker tongs
- f. funnel

\_\_\_\_ 35. This burn occurs when the skin or a mucous membrane is damaged by contact with a substance

\_\_\_\_ 36. NaOH



\_\_\_\_ 37.

\_\_\_\_ 38. This burn can occur if you touch a hot object or flame.



\_\_\_\_ 39.



\_\_\_\_ 40.

**Fall Semester Review Practice Test  
Answer Section****MULTIPLE CHOICE**

1. ANS: A                   PTS: 1
2. ANS: D                   PTS: 1
3. ANS: C                   PTS: 1
4. ANS: C                   PTS: 1
5. ANS: C                   PTS: 1
6. ANS: C                   PTS: 1
7. ANS: B                   PTS: 1
8. ANS: C                   PTS: 1
9. ANS: C                   PTS: 1
10. ANS: A                   PTS: 1
11. ANS: D                   PTS: 1
12. ANS: A                   PTS: 1
13. ANS: B                   PTS: 1
14. ANS: B                   PTS: 1
15. ANS: D                   PTS: 1
16. ANS: A                   PTS: 1
17. ANS: C                   PTS: 1
18. ANS: C                   PTS: 1
19. ANS: B                   PTS: 1
20. ANS: B                   PTS: 1
21. ANS: C                   PTS: 1
22. ANS: D                   PTS: 1
23. ANS: C                   PTS: 1
24. ANS: D                   PTS: 1
25. ANS: B                   PTS: 1
26. ANS: A                   PTS: 1

**MULTIPLE RESPONSE**

27. ANS: A, B               PTS: 1
28. ANS: C, D, E           PTS: 1
29. ANS: B, D               PTS: 1
30. ANS: A, B, C           PTS: 1
31. ANS: A, B, D           PTS: 1
32. ANS: A, C, D           PTS: 1
33. ANS: B, E, F           PTS: 1
34. ANS: A, C               PTS: 1



**MATCHING**

- |            |        |
|------------|--------|
| 35. ANS: A | PTS: 1 |
| 36. ANS: B | PTS: 1 |
| 37. ANS: E | PTS: 1 |
| 38. ANS: C | PTS: 1 |
| 39. ANS: F | PTS: 1 |
| 40. ANS: D | PTS: 1 |