Polyatomic Ions

Make flashcards for the following polyatomic ions.

Example Flashcard:

Front of Card

$C_2H_3O_2^{-1}$

Back of Card

Acetate

Charge = $^{-}$1

Formula	*Name*
1. $H_2PO_4^-$	Dihydrogen phosphate
2. $C_2H_3O_2^-$	Acetate
3. HSO_3^-	Hydrogen sulfite
4. HSO_4^-	Hydrogen sulfate
5. HCO_3^-	Hydrogen carbonate
6. NO_2^-	Nitrite
7. NO_3^-	Nitrate
8. CN^-	Cyanide
9. OH^-	Hydroxide
10. MnO_4^-	Permanganate
11. ClO^-	Hypochlorite
12. ClO_2^-	Chlorite
13. ClO_3^-	Chlorate
14. ClO_4^-	Perchlorate

Charge = -2

Formula	*Name*
15. HPO_4^{-2}	Hydrogen phosphate
16. $C_2O_4^{-2}$	Oxalate
17. SO_3^{-2}	Sulfite
18. SO_4^{-2}	Sulfate
19. CO_3^{-2}	Carbonate
20. CrO_4^{-2}	Chromate
21. $Cr_2O_7^{-2}$	Dichromate
22. SiO_3^{-2}	Silicate

Charge = -3

Formula	*Name*
23. PO_3^{-3}	Phosphite
24. PO_4^{-3}	Phosphate

Charge = +1

Formula	*Name*
25. NH_4^+	Ammonium