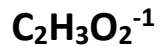


Polyatomic Ions

Make flashcards for the following polyatomic ions.

Example Flashcard:

Front of Card



Back of Card

Acetate

Charge = -1

| <i>Formula</i> | <i>Name</i> |
|---------------------------------------|----------------------|
| 1. H_2PO_4^- | Dihydrogen phosphate |
| 2. $\text{C}_2\text{H}_3\text{O}_2^-$ | Acetate |
| 3. HSO_3^- | Hydrogen sulfite |
| 4. HSO_4^- | Hydrogen sulfate |
| 5. HCO_3^- | Hydrogen carbonate |
| 6. NO_2^- | Nitrite |
| 7. NO_3^- | Nitrate |
| 8. CN^- | Cyanide |
| 9. OH^- | Hydroxide |
| 10. MnO_4^- | Permanganate |
| 11. ClO^- | Hypochlorite |
| 12. ClO_2^- | Chlorite |
| 13. ClO_3^- | Chlorate |
| 14. ClO_4^- | Perchlorate |

Charge = -2

| <i>Formula</i> | <i>Name</i> |
|----------------------------------|--------------------|
| 15. HPO_4^{2-} | Hydrogen phosphate |
| 16. $\text{C}_2\text{O}_4^{2-}$ | Oxalate |
| 17. SO_3^{2-} | Sulfite |
| 18. SO_4^{2-} | Sulfate |
| 19. CO_3^{2-} | Carbonate |
| 20. CrO_4^{2-} | Chromate |
| 21. $\text{Cr}_2\text{O}_7^{2-}$ | Dichromate |
| 22. SiO_3^{2-} | Silicate |

Charge = -3

| <i>Formula</i> | <i>Name</i> |
|------------------------|-------------|
| 23. PO_3^{3-} | Phosphite |
| 24. PO_4^{3-} | Phosphate |

Charge = +1

| <i>Formula</i> | <i>Name</i> |
|---------------------|-------------|
| 25. NH_4^+ | Ammonium |