

# Molecular Naming Examples

IS4T1

# IS4T1 – Naming Molecules

number of atoms	prefix	Chemical Formula	Chemical Name
1	mono	NO	<b>nitrogen monoxide</b>
2	di	NO <sub>2</sub>	<b>nitrogen dioxide</b>
3	tri	N <sub>2</sub> O <sub>3</sub>	<b>dinitrogen trioxide</b>
4	tetra	N <sub>2</sub> O <sub>4</sub>	<b>dinitrogen tetroxide</b>
5	<u>penta</u>	N <sub>2</sub> O <sub>5</sub>	<b>dinitrogen pentoxide</b>
6	<u>hexa</u>	SF <sub>6</sub>	<b>sulfur hexafluoride</b>
7	<u>hepta</u>	IF <sub>7</sub>	<b>iodine heptafluoride</b>
8	<u>octa</u>	P <sub>4</sub> O <sub>8</sub>	<b>tetraphosphorus octoxide</b>
9	<u>nona</u>	P <sub>4</sub> S <sub>9</sub>	<b>tetraphosphorus nonasulfide</b>
10	<u>deca</u>	As <sub>4</sub> O <sub>10</sub>	<b>tetraarsenic decoxide</b>

1) Using the guidelines above name or give formulas for these covalent substances:

a)  $\text{CO}_2$  carbon dioxide

b)  $\text{NBr}_3$  nitrogen tribromide

c) phosphorous tribromide  $\text{PBr}_3$

d)  $\text{CO}$  carbon monoxide

e)  $\text{SCl}_6$  sulfur hexachloride

f) dihydrogen monoxide  $\text{H}_2\text{O}$

g)  $\text{B}_2\text{H}_6$  diboron hexahydride